

Exercise-1

🐞 Marked Questions may have for Revision Questions.

ONLY ONE OPTION CORRECT TYPE**Section (A) : Oxidation number**

- A-1.** The oxidation number that iron does not exhibit in its common compounds or in its elemental state is -
(A) 0 (B) +1 (C) +2 (D) +3
- A-2. 🐞** The oxidation state of nitrogen varies from:
(A) -3 to +5 (B) 0 to +5 (C) -3 to 1 (D) +3 to +5
- A-3.** When H_2SO_3 is converted into H_2SO_4 the change in the oxidation state of sulphur is from :
(A) 0 to +2 (B) +2 to +4 (C) +4 to +2 (D) +4 to +6
- A-4.** The halogen that shows same oxidation state in all its compounds with other elements is :
(A) I_2 (B) F_2 (C) Cl_2 (D) Br_2
- A-5.** Most stable oxidation state of gold is -
(A) +1 (B) +3 (C) +2 (D) zero
- A-6. 🐞** The most stable oxidation state of chromium is -
(A) +5 (B) +3 (C) +2 (D) +4
- A-7.** Which can have both +ve and -ve oxidation states?
(A) F (B) I (C) Na (D) He
- A-8.** Which metal exhibits more than one oxidation states?
(A) Na (B) Mg (C) Al (D) Fe
- A-9.** The most common oxidation state of an element is -2. The number of electrons present in its outer most shell is :
(A) 2 (B) 4 (C) 6 (D) 8
- A-10.** Conversion of PbSO_4 to PbS the oxidation number of sulphur in PbS is-
(A) -2 (B) +6 (C) +4 (D) -1
- A-11. 🐞** Oxidation state of oxygen in H_2O_2 is :
(A) -2 (B) -1 (C) +1 (D) +2
- A-12.** The oxidation number of phosphorus in $\text{Mg}_2\text{P}_2\text{O}_7$ is :
(A) +5 (B) -5 (C) +6 (D) -7
- A-13.** In the conversion of Br_2 to BrO_3^- , the oxidation state of bromine changes from-
(A) 0 to +5 (B) -1 to +5 (C) 0 to -3 (D) +2 to +5
- A-14.** Oxidation number of S in S_2Cl_2 is :
(A) +1 (B) +6 (C) 0 (D) -1

A-15. Which of the following element shows only -1 oxidation number in combined state :

- (A) F (B) Cl (C) Br (D) I

A-16. The oxidation number of Fe in FeS_2 is

- (A) $+4$ (B) $+2$ (C) $+1$ (D) zero

Section (B) : Inorganic nomenclature

B-1. Correct formula of aluminium perchlorate is :

- (A) $\text{Al}(\text{ClO})_3$ (B) $\text{Al}(\text{ClO}_2)_3$ (C) $\text{Al}_2(\text{ClO}_3)_3$ (D) $\text{Al}(\text{ClO}_4)_3$

B-2. Sodium chlorite is :

- (A) NaClO_3 (B) NaClO_2 (C) NaClO (D) NaClO_4

B-3. Aluminium phosphide is :

- (A) AlP_3 (B) Al_2P_3 (C) AlP (D) Al_3P_2

B-4. Formula of Dioxygen difluoride is :

- (A) OF_2 (B) O_2F (C) O_2F_2 (D) O_2F_3

B-5. Barium azide is :

- (A) BaN (B) Ba_2N_3 (C) $\text{Ba}(\text{N}_3)_2$ (D) Ba_3N_2

B-6. Silicon fluoride Formula is :

- (A) SiF (B) SiF_3 (C) SiF_4 (D) SiF_6

B-7. Aluminium carbide is :

- (A) Al_2C (B) Al_4C_3 (C) AlC_3 (D) AlC

B-8. Which of the following oxyacids forms pyroxyacids :

- (A) H_3PO_4 (B) H_3BO_3 (C) H_2SO_4 (D) All of these

B-9. Which of the following set of element not forms metaoxy acids:

- (A) Cl, S, N (B) Cl, S, P (C) Si, C, B (D) C, Si, P

B-10. Name of oxyanion of boric acid (H_3BO_3) is :

- (A) Borate ion (B) Boraite ion (C) Hypo Borite ion (D) Per borate ion

B-11. Correct match is :

- (i) CrO_4^{2-} = chromate (ii) MnO_4^{2-} = Mangnate (iii) BO_3^{3-} = Borate (iv) XeO_4^{2-} = Xenate
(A) Only (i) (ii) (B) Only (ii) (iii) (C) Only (iii) (iv) (D) All of these

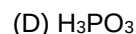
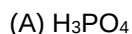
B-12. Sodium tri-sulphide Formula is :

- (A) Na_2S_3 (B) Na_3S (C) Na_3S_2 (D) Na_2S

B-13. PO_4^{3-} is :

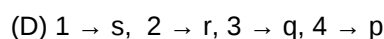
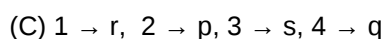
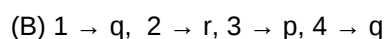
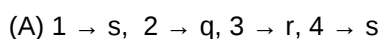
- (A) Phosphate ion (B) Phasphite ion (C) Hypophosphite ion (D) Pyrophosphite ion

B-14. Pyrophosphoric acid is :



B-15. Correctly match codes are :

(1)	H_3PO_4	(p)	Meta phosphoric acid
(2)	HPO_3	(q)	Thio sulphuric acid
(3)	H_2SO_4	(r)	Phosphoric acid
(4)	$\text{H}_2\text{S}_2\text{O}_3$	(s)	sulphuric acid



Exercise-2

Marked Questions may have for Revision Questions.

ONLY ONE OPTION CORRECT TYPE

- The oxidation number of sulphur in H_2SO_5
(A) +3 (B) +5 (C) +6 (D) +8
- Which of the following is a correct statement.
(A) The name of SeO_3^{2-} is selenite
(B) CdF_3 is correct formula
(C) Zn can only show the oxidation state +3
(D) The element having atomic numbers 29 lies in p-block
- In the given formula which one is correct :
(A) SiI_2 (B) Cr_2O_7 (C) Ti_2O_5 (D) $\text{Na}_2\text{C}_8\text{H}_4\text{O}_4$
- Prefix per- is attached to the name
(A) H_2SnO_3 (B) Sb_2O_5 (C) H_3PO_5 (D) HNO_2
- Match column-I with column II and select correct.

	Column-I		Column-II
(I)	CO_3^{2-}	(P)	Carbonate ion
(II)	N_3^-	(Q)	Azide ion
(III)	O_2^{2-}	(R)	Acetate ion
(IV)	CH_3COO^-	(S)	Peroxide ion

Code :

	I	II	III	IV		I	II	III	IV
(A)	P	Q	R	S	(B)	P	Q	S	R
(C)	R	S	Q	R	(D)	R	P	Q	S

- Dichromate ion is :
(A) CrO_4^{2-} (B) $\text{Cr}_2\text{O}_7^{2-}$ (C) CrO_3 (D) Cr_2O_4

7. In following compound dithionic acid is :
 (A) $\text{H}_2\text{S}_2\text{O}_6$ (B) $\text{H}_2\text{S}_2\text{O}_4$ (C) H_2SO_5 (D) $\text{H}_2\text{S}_2\text{O}_3$
8. Strontium metaphosphate is
 (A) $\text{Sr}(\text{PO}_3)_2$ (B) SrHPO_3 (C) $\text{Sr}_3(\text{PO}_4)_2$ (D) $\text{Sr}_2\text{P}_2\text{O}_7$
9. Nickel (II) pyroselenate is
 (A) $\text{Ni}_2\text{Se}_2\text{O}_7$ (B) NiSe_2O_7 (C) $\text{Ni}_2\text{Se}_2\text{O}_5$ (D) NiSe_2O_5
10. The formula of sodium tungstate is Na_2WO_4 and that of lead phosphate is $\text{Pb}_3(\text{PO}_4)_2$. What is the formula for lead tungstate ?
 (A) PbWO_4 (B) $\text{Pb}_2(\text{WO}_4)_3$ (C) $\text{Pb}_3(\text{WO}_4)_2$ (D) $\text{Pb}_3(\text{WO}_4)_4$
11. Mercurous azide is
 (A) $\text{Hg}_2(\text{N}_3)_2$ (B) HgN_3 (C) Hg_2N_3 (D) $\text{Hg}(\text{N}_3)_2$
12. $\text{Fe}[\text{Fe}(\text{CN})_6]$ is
 (A) ferroferrocyanide (B) Ferriferrocyanide (C) ferroferricyanide (D) ferriferrocyanide
13. Ethyl methyl ether, $\text{CH}_3\text{—O—C}_2\text{H}_5$, is used as an anaesthetic. Formula for corresponding thioether would be
 (A) $\text{CH}_3\text{—S—C}_2\text{H}_5$ (B) $\text{CH}_3\text{—O—S—C}_2\text{H}_5$ (C) $\text{C}_2\text{H}_5\text{—O—CH}_3$ (D) $\text{C}_2\text{H}_5\text{—O—CH}_2\text{SH}$
14. Hydracid which contains nitrogen is
 (A) HN_3 (B) HNO_3 (C) HNO_2 (D) NH_3
15. Anhydride of HClO_4 is
 (A) Cl_2O_7 (B) ClO_3 (C) Cl_2O_5 (D) ClO_2
16. Correct name of the compound NaCrO_2 will be
 (A) Sodium metachromate (B) Sodium metachromite
 (C) Sodium orthochromate (D) Sodium orthochromite
17. A-pyro acid cannot be formed by
 (A) HClO_4 (B) H_2SO_4 (C) H_3PO_4 (D) H_3BO_3
18. Correct name for $\text{Na}_2\text{CaP}_2\text{O}_7$ is
 (A) Sodium calcium pyrophosphate (B) Sodium calcium metaphosphate
 (C) Sodium calcium orthophosphate (D) None of these
19. Correct formula for rubidium metagallate is
 (A) RbGaO_2 (B) Ru_2GeO_2 (C) Rb_3GaO_3 (D) Ru_2GaO_3

Answers**EXERCISE - 1**

A-1. (B)	A-2. (A)	A-3. (D)	A-4. (B)	A-5. (D)
A-6. (B)	A-7. (B)	A-8. (D)	A-9. (C)	A-10. (A)
A-11. (B)	A-12. (A)	A-13. (A)	A-14. (A)	A-15. (A)
A-16. (B)	B-1. (D)	B-2. (B)	B-3. (C)	B-4. (C)
B-5. (C)	B-6. (C)	B-7. (B)	B-8. (D)	B-9. (A)
B-10. (A)	B-11. (D)	B-12. (A)	B-13. (A)	B-14. (C)
B-15. (C)				

EXERCISE - 2

1. (C)	2. (A)	3. (D)	4. (C)	5. (B)
6. (B)	7. (A)	8. (A)	9. (B)	10. (A)
11. (A)	12. (B)	13. (A)	14. (A)	15. (A)
16. (B)	17. (A)	18. (A)	19. (A)	