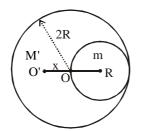


# HINTS & SOLUTIONS

2.

### SECTION I - PHYSICS

1. (b) Let O be the centre of mass of the disc having radius 2R. O' is the new C.M.



Let m = mass of disc of radius R M' = mass of disc when the disc of radius R is removed. M = mass of disc of radius 2RNow,  $m = (\pi R^2) . \sigma$ , where  $\sigma = \frac{M}{\pi (2R)^2} = \frac{M}{4\pi R^2}$  = the mass per unit area  $M' = [\pi(2R)^2 - \pi R^2].\sigma$ 4.  $= 3\pi R^2 \sigma$  $M = \pi (2R)^2 . \sigma = 4\pi R^2 \sigma$ M'.x + m∕₽ We have, (:: C.M. of the full disc is at the centre O)M'.x + mRor, M'x or,  $\left(\frac{2\pi R^2 \sigma}{2\pi P^2 \sigma}\right) R = \left(-\frac{1}{3}\right) R$  $3\pi R^2 \sigma$ R 5.  $\frac{\alpha}{R}$ 

There appears misprint in this question.

There must be  $\alpha R$  instead of  $\mathcal{F}$ . Then

$$\alpha R = \left(-\frac{1}{3}\right) R = \left(\alpha = \frac{1}{3}\right)$$

$$\therefore \quad |\alpha| = \frac{1}{3}$$

(b) The acceleration of a solid sphere of mass M, radius t and moment of inertia I rolling down (without slipping) an inclined plane making an angle  $\theta$  with the horizontal is given by

 $\underbrace{g\sin\theta}_{I+\frac{K^2}{R^2}}, \text{ where, } I = MK^2$ 

- Central forces always act along the axis of rotation. Therefore, the torque is zero. And if there is no external torque acting on a rotating body then its angular momentum is constant.
- (b) Let the spring be compressed by x.
   Clearly, Initial K.E. of block = Potential energy of spring + workdown against friction

or, 
$$\frac{1}{2}mv^2 = \frac{1}{2}kx^2 + fx$$
  
or,  $\frac{1}{2} \times 2 \times (4)^2 = \left(\frac{1}{2} \times 10000 \times x^2\right) + 15x$   
or.  $16 = 5000x^2 + 15x$ 

or, 
$$5000 x^2 + 15x - 16 = 0$$

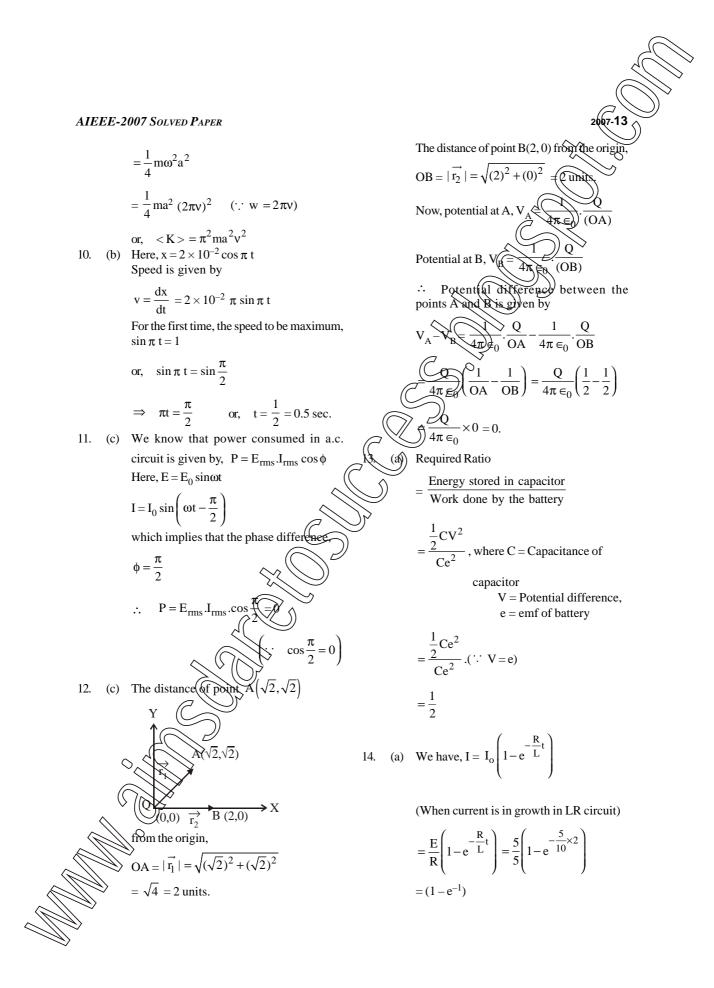
$$\therefore \quad x = \frac{-15 \pm \sqrt{(15)^2 - 4 \times 5000 \times (-16)}}{2 \times 5000}$$
$$= \frac{-15 \pm 565.88}{10000} = 0.055 \,\mathrm{m}$$

(Ignoring -ve value)

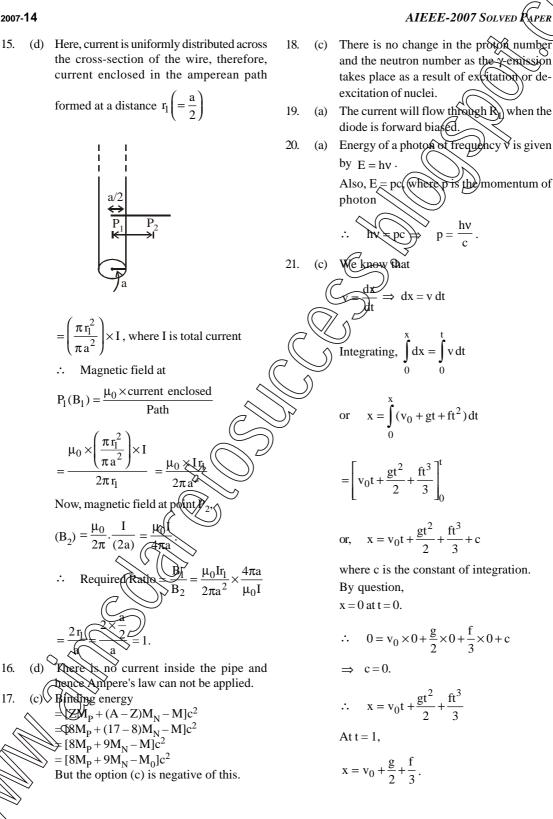
 $\therefore x = 5.5 \text{ cm.}$ (d) Let K' be the K.E. at the highest point. Then

 $\mathbf{K}' = \frac{1}{2} \mathbf{m} \mathbf{v}_x^2 (\because \mathbf{v}_y = 0 \text{ at highest point})$ 

 $\sim$ 

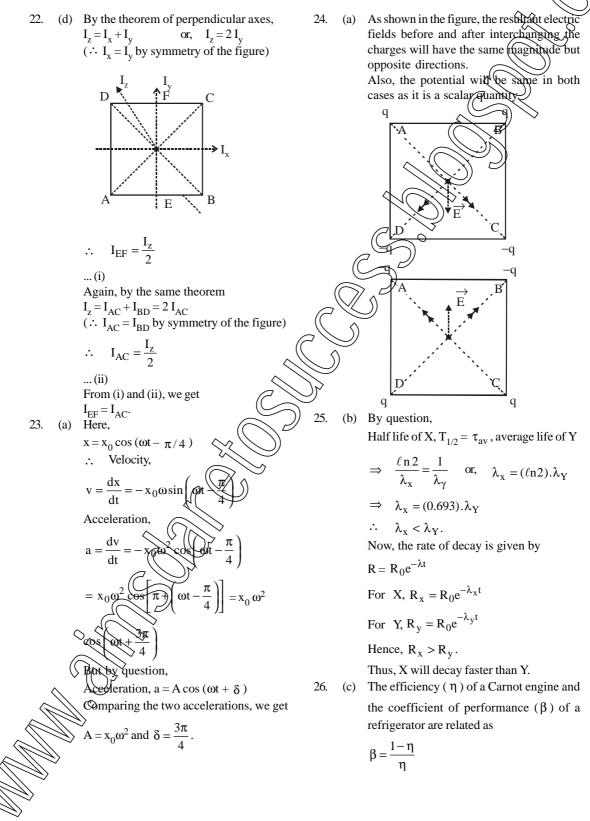






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29.

Here, 
$$\eta = \frac{1}{10}$$
  
 $\therefore \quad \beta = \frac{1 - \frac{1}{10}}{\left(\frac{1}{10}\right)} = 9.$ 

Also, Coefficient of performance  $(\beta)$  is

given by 
$$\beta = \frac{Q_2}{w}$$
, where  $Q_2$  is the energy absorbed from the reservoir.

or, 
$$9 = \frac{Q_2}{10}$$

$$\therefore Q_2 = 90 \text{ J.}$$

- 27. (a) Si and Ge are semiconductors but C is an insulator. Also, the conductivity of Si and Ge is more than C because the valence electrons of Si, Ge and C lie in third, fouth and second orbit repsectively.
- 28. (b) Here,  $\vec{E}$  and  $\vec{B}$  are perpendicular to each ( other and the velocity  $\vec{v}$  does not change therefore

$$qE = qvB \implies v = \frac{E}{B}$$
Also,  $\left|\frac{\vec{E} \times \vec{B}}{B^2}\right| = \frac{E B \sin \theta}{B^2}$ 

$$= \frac{E B \sin 90^\circ}{B^2} = \frac{E}{B} = |\vec{v}| = v$$

$$\therefore \text{ Option (b) is correct.}$$
(a) Here,  $V(x) = (\vec{v} + 4)$ 
We know that  $E = -\frac{dv}{dx} = -\frac{d}{dx} \left(\frac{20}{x^2 - 4}\right)$ 

$$40x$$

 $\frac{160}{144} = +\frac{10}{9}$  volt /  $\mu$ m.

Positive sign indicates that  $\vec{E}$  is in +ve x-

direction.

30. (d) We have to find the frequency of emitted photons. For emission of photons the transition must take place from a higher energy level to a lower energy level which are given only in options (c) and (d). Frequency is given by 7

$$hv = -13.6 \left( \frac{1}{n_1^2} + \frac{1}{n_2^2} \right)$$
  
For transition from  $n = 6$  to  $n = 2$ ,  
 $v_1 = -\frac{13.6}{n} + \frac{1}{2^2} = \frac{2}{9} \times \left( \frac{13.6}{n} \right)$   
For transition from  $n = 2$  to  $n = 1$ 

$$\int_{2}^{2} \frac{13.8}{2} \left(\frac{1}{2^2} - \frac{1}{1^2}\right) = \frac{3}{4} \times \left(\frac{13.6}{h}\right)$$

we get T = ma and F - T = Mawhere T is force due to spring  $\Rightarrow F - ma = Ma$ or, F = Ma + ma

$$\therefore$$
  $a = \frac{F}{M+m}$ .

Now, force acting on the block of mass m is

$$ma = m\left(\frac{F}{M+m}\right) = \frac{mF}{m+M}$$

32. (c) Power of combination is given by  $P=P_1+P_2=(-15+5)D=-10D.$ 

Now, 
$$P = \frac{1}{f} \implies f = \frac{1}{P} = \frac{1}{-10}$$
 metre  
 $\therefore f = -\left(\frac{1}{10} \times 100\right)$  cm = -10 cm.

35.

36.

(c)

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33. (d) Let T be the temperature of the interface. As the two sections are in series, the rate of flow of heat in them will be equal.

$$\therefore \quad \frac{\mathrm{K}_{1}\mathrm{A}(\mathrm{T}_{1}-\mathrm{T})}{\ell_{1}} = \frac{\mathrm{K}_{2}\mathrm{A}(\mathrm{T}-\mathrm{T}_{2})}{\ell_{2}},$$

where A is the area of cross-section.

or, 
$$K_1 A (T_1 - T) \ell_2 = K_2 A (T - T_2) \ell_1$$
  
or  $K_1 T_1 \ell_2 = K_2 T \ell_2 = K_2 T \ell_2 = K_2 T_2$ 

or, 
$$K_1T_1\ell_2 - K_1T\ell_2 = K_2T\ell_1 - K_2T_2\ell_1$$
  
or,  $(K_2\ell_1 + K_1\ell_2)T = K_1T_1\ell_2 + K_2T_2\ell_1$ 

or, 
$$(\mathbf{K}_2 \iota_1 + \mathbf{K}_1 \iota_2)\mathbf{I} = \mathbf{K}_1 \mathbf{I}_1 \iota_2 + \mathbf{K}_2 \mathbf{I}_2 \iota$$

$$\therefore \quad \mathbf{T} = \frac{\mathbf{K}_{1}\mathbf{T}_{1}\ell_{2} + \mathbf{K}_{2}\mathbf{T}_{2}\ell_{1}}{\mathbf{K}_{2}\ell_{1} + \mathbf{K}_{1}\ell_{2}}$$
$$= \frac{\mathbf{K}_{1}\ell_{2}\mathbf{T}_{1} + \mathbf{K}_{2}\ell_{1}\mathbf{T}_{2}}{\mathbf{K}_{1}\ell_{2} + \mathbf{K}_{2}\ell_{1}\mathbf{T}_{2}}.$$

34. (a) We have, 
$$L_1 = 10 \log \left( \frac{I_1}{I_0} \right)$$

$$L_{2} = 10 \log \left( \frac{I_{2}}{I_{0}} \right)$$

$$\therefore \quad L_{1} - L_{2} = 10 \log \left( \frac{I_{1}}{I_{0}} - 10 \log \frac{I_{2}}{I_{0}} \right)$$
or,  $\Delta L = 10 \log \left( \frac{I_{1}}{I_{0}} - 10 \log \frac{I_{2}}{I_{0}} \right)$ 
or,  $\Delta L = 10 \log \left( \frac{I_{1}}{I_{2}} - 10 \log \frac{I_{1}}{I_{2}} \right)$ 
or,  $\Delta L = 10 \log \left( \frac{I_{1}}{I_{2}} - 10 \log \frac{I_{1}}{I_{2}} \right)$ 
or,  $\Delta L = 10 \log \left( \frac{I_{1}}{I_{2}} - 10 \log \frac{I_{1}}{I_{2}} \right)$ 
or,  $I_{2} = 10^{2}$ 
or,  $I_{2} = \frac{I_{1}}{100}$ .
$$\Rightarrow \text{ Intensity decreases by a factor 100.}$$

(b) We have, Molar heat capacity = Molar mass pec heat capacity per unit mass  $C_n = 28 C_p$  (for introgen *:*.. and  $C_v^P = 28 C$ Now,  $C_p - C_v =$ 

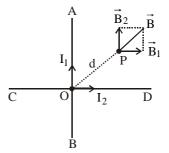
=R

or, 
$$28C_p - 28C_p = R = R$$
  
 $\Rightarrow C_p - C_p = R$   
When a charged particle er  
field at a direction perpe

f is given by  $\frac{1}{2}$  mv<sup>2</sup> and v<sup>2</sup> is the square

of the magnitude of velocity which does not change.

Clearly, the magnetic fields at a point P, equidistant from AOB and COD will have directions perpendicular to each other, as they are placed normal to each other.



Resultant field,  $B = \sqrt{B_1^2 + B_2^2}$ :.

But 
$$B_1 = \frac{\mu_0 I_1}{2\pi d}$$
 and  $B_2 = \frac{\mu_0 I_2}{2\pi d}$ 

$$\therefore \qquad \mathbf{B} = \sqrt{\left(\frac{\mu_0}{2\pi d}\right)^2 \left(\mathbf{I}_1^2 + \mathbf{I}_2^2\right)}$$

or, 
$$B = \frac{\mu_0}{2\pi d} (I_1^2 + I_2^2)^{1/2}$$

4

where

### 2007-18

38. (d) We know that

$$R_t = R_0 (1 + \alpha t),$$

$$\Rightarrow R_{50} = R_0 (1 + 50 \alpha) \qquad \dots (i)$$
  
$$R_{100} = R_0 (1 + 100 \alpha) \qquad \dots (ii)$$

From (i), 
$$R_{50} - R_0 = 50 \alpha R_0$$
 ... (iii)

From (ii), 
$$R_{100} - R_0 = 100 \alpha R_0$$
 ... (iv)

From (ii),  $R_{100} - R_0 = 100 \,\alpha R_0$ Dividing (iii) by (iv), we get

$$\frac{R_{50} - R_0}{R_{100} - R_0} = \frac{1}{2}$$

Here,  $R_{50} = 5\Omega$  and  $R_{100} = 6\Omega$ 

$$\therefore \quad \frac{5 - R_0}{6 - R_0} = \frac{1}{2}$$
  
or, 
$$6 - R_0 = 10 - 2$$

or, 
$$R_0 = 4\Omega$$
.

39. The potential energy of a charged capacitor (a)

is given by 
$$U = \frac{Q^2}{2C}$$

If a dielectric slab is inserted between the

 $R_0$ 

plates, the energy is given by 2KC

K is the dielectric constant. Again, when the dielectric state is removed slowly its energy increases to initial potential energy. Thus, work done is zero.

(b) Electronic charge does not depend on 40. acceleration due to gravity as it is a universal constant. So, electronic charge on earth = electronic charge on moon Required ratio 1 ÷.

# SECTION II - CHEMISTRY

41. (b) Occording to Kohlrausch's law, molar conductivity of weak electrolyte acetic acid (CH3COOH) is given as follows:

 $\Lambda^{\circ}_{CH_{3}COOH} = \Lambda^{\circ}_{CH_{3}COONa} + \Lambda^{\circ}_{HCl} - \Lambda^{\circ}_{NaCl}$ 

Value of  $\Lambda^{\circ}_{NaCl}$  should also be

known for calculating value of  $\Lambda^{\circ}_{CH_3COOH}$ .

42. (d) Aromatic amines are less basic than aliphatic  
amines. Among aliphatic amines the order of  
basicity is 
$$2^{\circ} > 1^{\circ} > 3^{\circ}$$
 (:: of decreased  
electron density due to reording in 3°  
amines)  
 $\therefore$  dimethylamine (27 aliphatic amine) is  
strongest base among given hoices.  
43. (d) When alkyl benzend are oxidized with alkaline  
KMTD<sub>4</sub>, the entire alkyl group is oxidised to  
-COOH group regardless of length of side  
chain.  
44. (a)  $(H_3 - CH_3 - CH_3 - 2 - 1)$   
 $(H_3 - CH_2 - CH_2 - C - CH - CH_2 - CH_3$   
 $(H_3 - CH_2 - CH_2 - C - CH - CH_2 - CH_3$   
 $(H_3 - CH_2 - CH_2 - C - CH - CH_2 - CH_3$ 

CH<sub>3</sub> 3-ethyl-4,4-dimethyl heptane

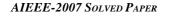
45. (b) Diamagnetic species have no unpaired electrons

$$O_2^{2-} \Rightarrow \sigma 1s^2, \ \sigma^* 1s^2, \ \sigma^* s^2, \ \sigma 2p_z^2, \ \pi 2p_x^2, \ \pi^2 2p_x^2, \ \pi^* 2p_x^2, \ \pi^* 2p_x^2, \ \pi^* 2p_x^2$$

- 46. (c) Reluctance of valence shell electrons to participate in bonding is called inert pair effect. The stability of lower oxidation state (+2 for group 14 element) increases on going down the group. So the correct order is  $SiX_2 < GeX_2 < PbX_2 < SnX_2$
- 47. (d) Chlorine reacts with excess of ammonia to produce ammonium chloride and nitrogen.

$$8NH_3 + 3Cl_2 \longrightarrow N_2 + NH_4Cl$$

- 48. (d) Smaller the size and higher the charge more will be polarising power of cation. So the correct order of polarising power is  $K^+\!<\!Ca^{2+}\!<\!Mg^{2+}\!<\!Be^{2+}$
- 49. (d) Mass of 3.6 moles of  $H_2SO_4$ = Moles  $\times$  Molecular mass  $= 3.6 \times 98 \text{ g} = 352.8 \text{ g}$ 
  - 1000 ml solution has 352.8 g of  $H_2SO_4$ *.*..



Given that 29 g of  $H_2SO_4$  is present in = 100 g of solution *:*. 352.8 g of  $H_2SO_4$  is present in

 $=\frac{100}{29}\times352.8$  g of solution

= 1216 g of solution

Density =  $\frac{\text{Mass}}{\text{Volume}} = \frac{1216}{1000} = 1.216 \text{ g/ml}$ = 1.22 g/ml

50. (d) 
$$H_2A \rightleftharpoons H^+ + HA^-$$

$$\therefore$$
  $K_1 = 1.0 \times 10^{-5} = \frac{[H^+][HA]}{[H_2A]}$ 

$$\mathrm{HA}^{-} \longrightarrow \mathrm{H}^{+} + \mathrm{A}^{-}$$

$$\therefore \quad K_2 = 5.0 \times 10^{-10} = \frac{[H^+][A^-]}{[HA^-]}$$

$$K = \frac{[H^+]^2 [A^2]}{[H_2 A]} = K_1 \times K_2$$
  
= (1.0 × 10<sup>-5</sup>) × (5 × 10<sup>-10</sup>) = 5 × 10<sup>-15</sup>  
51. (b) Given p\_A^0 = ?, p\_B^0 = 200 \text{mm}, x\_A = 0.6, x\_B = 1 - 0.6 = 0.4, P = 290

$$P = p_A + p_B = p_A^0 x_A + p_B^0 x_B$$
  
⇒ 290 =  $p_A^0 \times 0.6 + 200 \times 0.4$   
∴  $p_A^0 = 350 \text{ mm}$ 

52. (a) 
$$\Delta G^{\circ} = \Delta H^{\circ} - T\Delta S^{\circ}$$
  
For a spontaneous reaction  $\Delta G^{\circ}$ 

or 
$$\Delta H^{\circ} - TAS^{\circ} \rightarrow T > \frac{\Delta H^{\circ}}{\Delta S^{\circ}}$$
  
 $\Rightarrow T > \frac{179.3 \times 10^3}{160.2} > 1117.9K \approx 1118$ 

$$\Rightarrow \overbrace{160.2}^{1} > 1117.9 \text{K} \approx 1118 \text{K}$$

< 0

53. (a) 
$$AH_R = E_P - E_b = 180 - 200 = -20 \text{ kJ/mol}$$
  
The nearest correct answer given in choices  
may be obtained by neglecting sign.  
54. (d) E  $\Omega = 0$ : when cell is completely discharged.

$$E_{cell} = E_{cell}^{\circ} - \frac{0.059}{2} \log \left( \frac{\left[ Zn^{2+} \right]}{\left[ Cu^{2+} \right]} \right)$$

or 
$$0 = 1.1 - \frac{0.059}{2} \log \left[ \frac{Zn^{2+}}{Cu^{2+}} \right]$$
  
 $\log \left[ \frac{Zn^{2+}}{Cu^{2+}} \right] = \frac{2}{0.05} + \frac{2}{0.0$ 

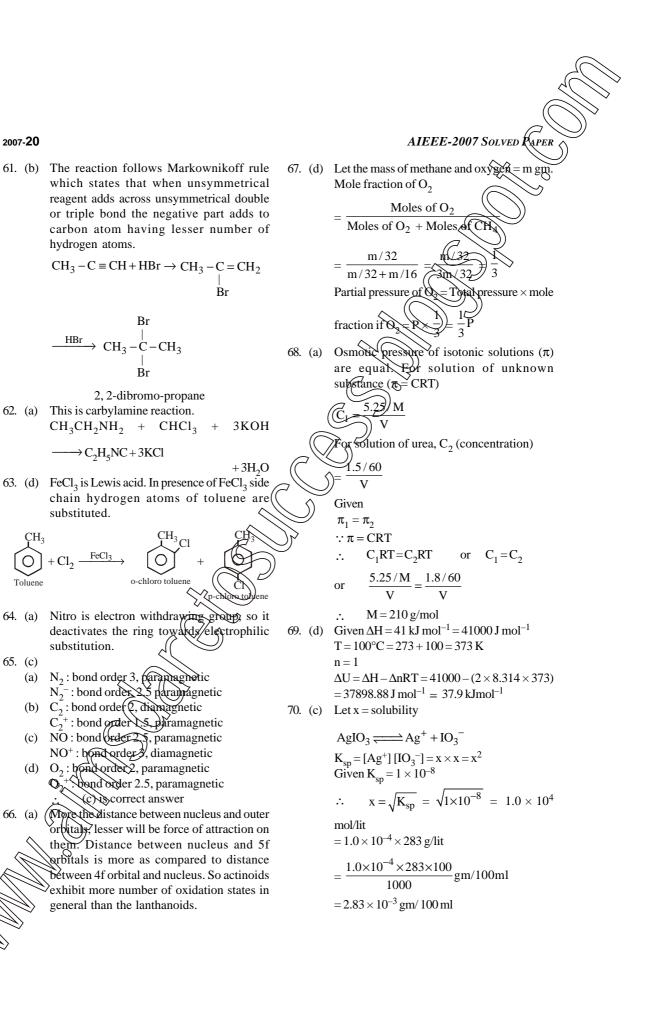
on with respect to  $\mathbf{B} = 1$  because change in concentration of B does not change half life. Order of reaction with respect to A = 1because rate of reaction doubles when concentration of A is doubled keeping concentration of A constant.

Order of reaction = 1 + 0 = 1 and units *:*. of first order reaction are  $L \mod^{-1} \sec^{-1}$ .

- 4f orbital is nearer to nucleus as compared to 57. (a) 5f orbital therefore, shielding of 4f is more than 5f.
- Complexes with  $dsp^2$  hybridisation are square planar. So  $[PtCl_4]^{2-}$  is square planar 58. (a) in shape.
- 59. (b) The organic compounds which have chiral carbon atom and do not have plane of symmetry rotate plane polarised light.

CHO HO - C - H(\* is asymmetric carbon) ĊH<sub>2</sub>OH

60. (b) Proteins have two types of secondary structures  $\alpha$ -helix and  $\beta$ -plated sheet.



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71. (a) Let activity of safe working = A Given  $A_0 = 10A$ 

$$\lambda = \frac{0.693}{t_{1/2}} = \frac{0.693}{30}$$
$$t_{1/2} = \frac{2.303}{\lambda} \log \frac{A_0}{A} = \frac{2.303}{0.693/30} \log \frac{10A}{A}$$
$$= \frac{2.303 \times 30}{0.693} \times \log 10 = 100 \text{ days.}$$

- 72. (b) Chiral conformation will not have plane of symmetry. Since twisted boat does not have plane of symmetry it is chiral.
- 73. (c) In  $S_N^2$  mechanism transition state is pentavelent. For bulky alkyl group it will have sterical hinderance and smaller alkyl group will favour the  $S_N^2$  mechanism. So the decreasing order of reactivity of alkyl halides is

 $\rightarrow CH_3CH_2MgI - (B)$ 

H<sub>2</sub>O

HCHO

(D)

propyl alcohol

$$RCH_2X > R_2CHX > R_3CX$$

74. (d)  $CH_3CH_2OH \xrightarrow{P+I_2} CH_3CH_2I$ 

Mg Ether

CH<sub>3</sub>CH<sub>2</sub>

H-C-OMgI

Ĥ

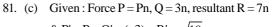
77. (c) 
$$2AI_{(s)}^{+}+6HCI_{(aq)}^{-} \rightarrow 2AI^{3+}_{(aq)}^{+}+6CI_{(aq)}^{+}+3H_{20}^{-}$$
  
 $\therefore 6 \text{ moles of HCl produces} = 8 \text{ moles of H}_{2}^{-}$   
 $\Rightarrow 22.41 \text{ of H}_{2}^{-}$   
 $\Rightarrow 1 \text{ mole of HCl produces}^{-} = 3 + 22.41 \text{ of H}_{2}^{-}$   
 $\Rightarrow 1 \text{ mole of HCl produces}^{-} = 3 + 22.41 \text{ of H}_{2}^{-}$   
 $\Rightarrow 2 \text{ moles of Al produces}^{-} = 3 \text{ moles of H}_{2}^{-}$   
 $\Rightarrow 2 \text{ moles of Al produces}^{-} = 3 \text{ moles of H}_{2}^{-}$   
 $\Rightarrow 1 \text{ mole of Al produces}^{-} = 33.6 \text{ L of H}_{2}^{-}$ 

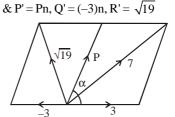
78. (a)  $(NH_4)_2SO_4 + 2H_2O \longrightarrow 2H_2SO_4 + NH_4OH H_2SO_4$  is strong acid and increases the activity of youl.

79. (b) Spontaneity of reaction depends on readency to acquire minimum energy state and maximum randomness. For a spontaneous process in an isolated system the change in entropy is positive.

(b)) Isotopes are atoms of same element having same atomic number but different atomic masses. Neutron has atomic number 0 and atomic mass 1. So loss of neutron will generate isotope.

# SECTION III - MATHEMATICS





We know that  $R^{2} = P^{2} + Q^{2} + 2PQ \cos \alpha$   $\Rightarrow (7)^{2} = P^{2} + (3)^{2} + 2 \times P \times 3 \cos \alpha$   $\Rightarrow 49 = P^{2} + 9 + 6P \cos \alpha \qquad \dots (i)$   $\Rightarrow 40 = P^{2} + 6P \cos \alpha$ and  $\left(\sqrt{19}\right)^{2} - P^{2} + (-3)^{2} + 2P \times -3 \cos \alpha$ 

and 
$$(\sqrt{19})^2 = P^2 + (-3)^2 + 2P \times -3 \cos \alpha$$

 $\Rightarrow 19 = P^2 + 9 - 6P \cos \alpha$   $\Rightarrow 10 = P^2 - 6P \cos \alpha \qquad \dots (ii)$ Adding (i) and (ii)  $50 = 2P^2$ 

$$\Rightarrow P^2 = 25 \Rightarrow P = 5n.$$

75. (c)

(a)  $n = 3, \ell \neq 0$  means 3s-orbital

(b) 
$$n =$$
 means 3p-orbital

(c) 
$$\mathbf{n} = \mathbf{3} + \mathbf{2}$$
 means 3d-orbital

(d) 
$$\sqrt{n} = 4, \ell = 0$$
 means 4s-orbital

ncreasing order of energy among these

 $\therefore$  3d has highest energy.

Greater the difference between electronegativity of bonded atoms, stronger will be bond.

 $\therefore$  F – H ..... F is the strongest bond.

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$$A = A_1$$

$$A_3 - A_1$$

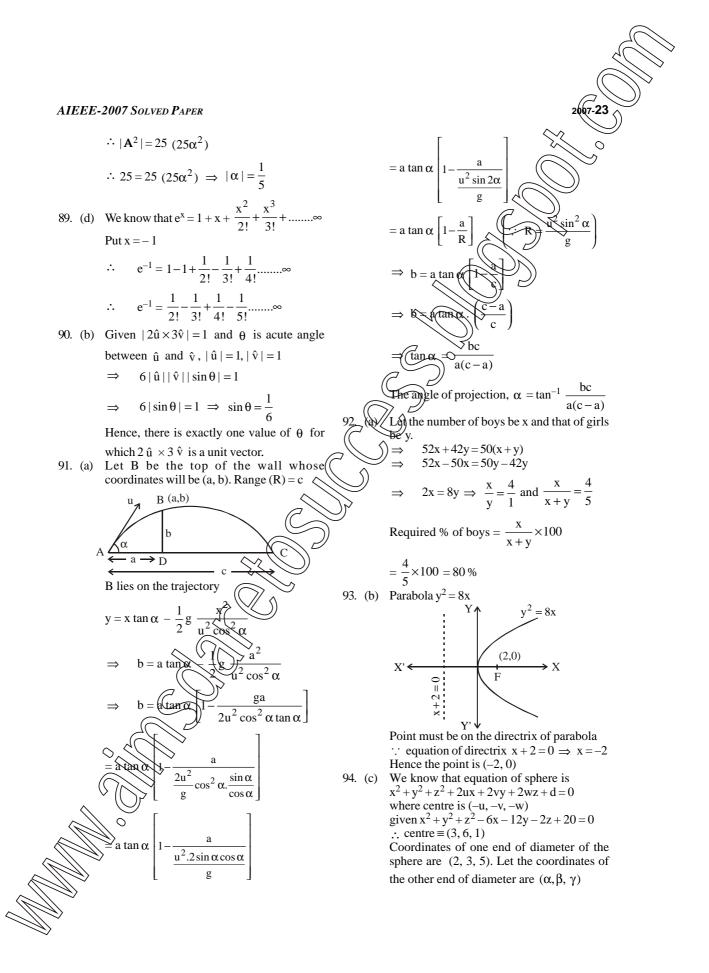
$$A = A$$$$$$

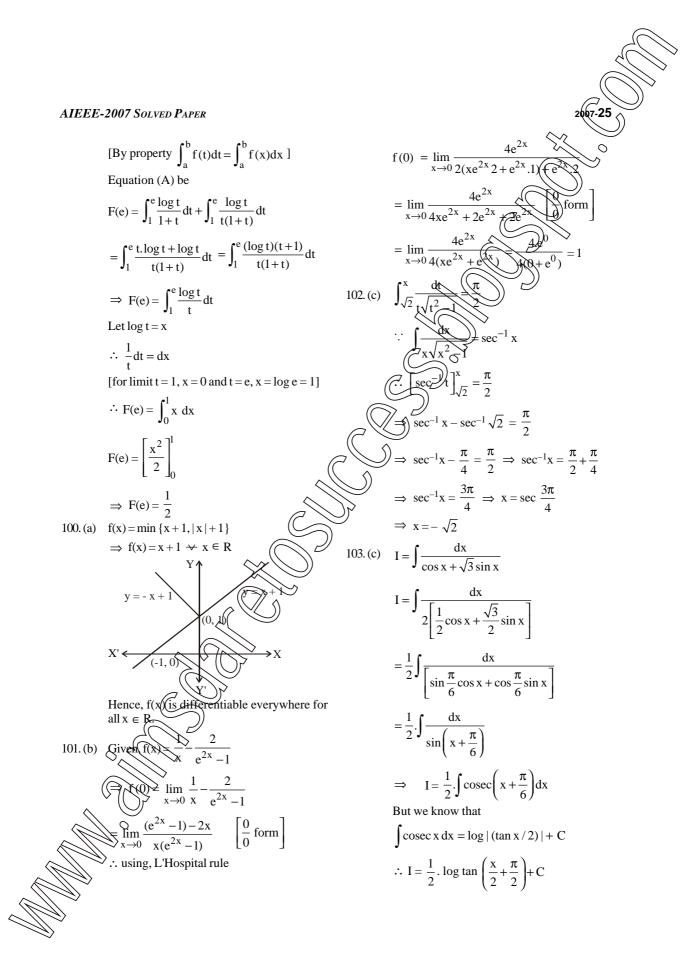
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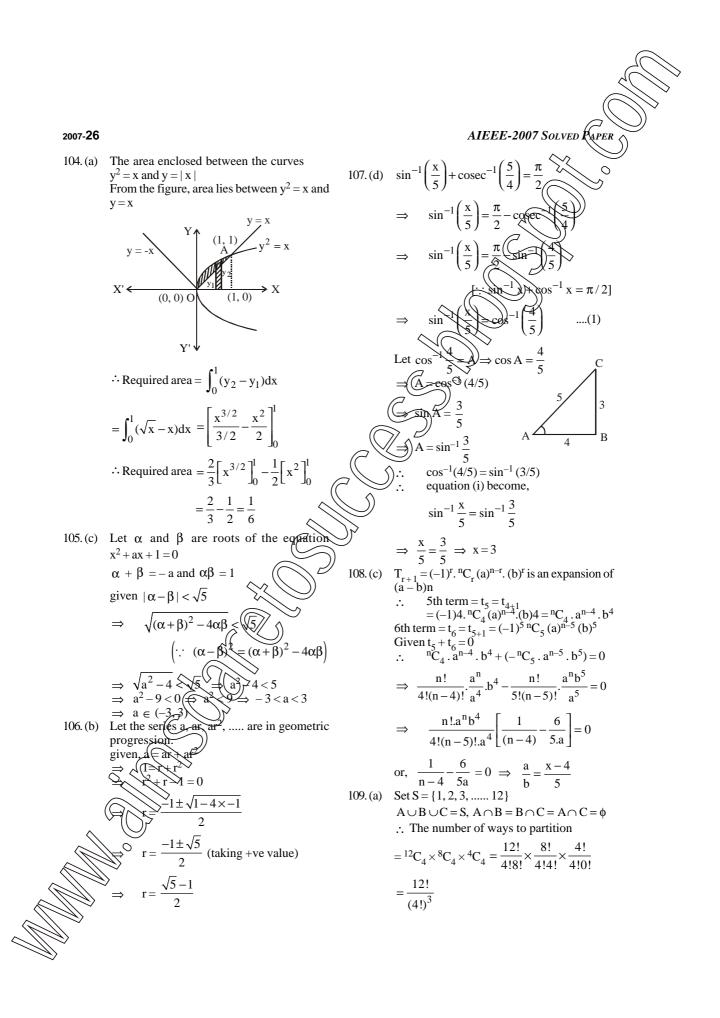
85. (b) Let the angle of line makes with the direction of z-axis is of direction c line with the +ve direction of x-axis and z-axis is 1, m, mespectively.

$$\therefore 1 = \cos \left( \frac{\pi}{4}, n = \cos \alpha \right)$$
  
as we know that  $t^{+} + m^{2} + n^{2} = 1$   
$$\Rightarrow \cos^{2} \frac{\pi}{4} + \cos^{2} \alpha = 1$$
  
$$\Rightarrow \cos^{2} \alpha = 0 \Rightarrow \alpha = \frac{\pi}{2}$$
  
Hence, angle with positive directio  
z-axis is  $\frac{\pi}{2}$ 

$$\Rightarrow \cos^2 \alpha = 0 \Rightarrow \alpha = \frac{\pi}{2}$$





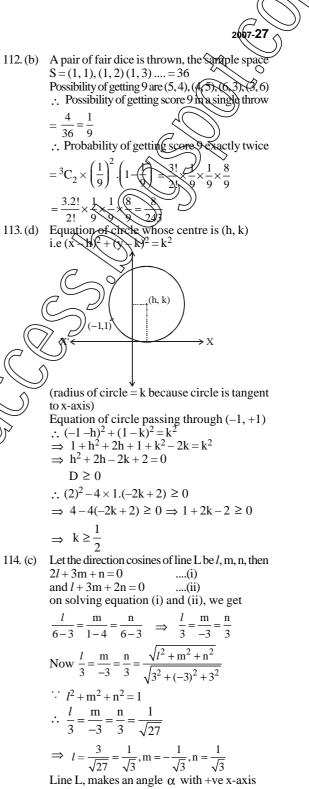


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110. (b) 
$$f(x) = 4^{-x^{2}} + \cos^{-1}\left(\frac{x}{2} - 1\right) + \log(\cos x)$$

$$f(x) \text{ is defined if } -1 \le \left(\frac{x}{2} - 1\right) \le 1 \text{ and } \cos x > 0$$
or
$$0 \le \frac{x}{2} \le 2 \text{ and } -\frac{\pi}{2} < x < \frac{\pi}{2}$$
or
$$0 \le x \le 4 \text{ and } -\frac{\pi}{2} < x < \frac{\pi}{2}$$

$$\therefore \quad x \in \left[0, \frac{\pi}{2}\right]$$
111. (a) Given : A body weighing 13 kg is suspended



 $\therefore l = \cos \alpha$ 

$$\Rightarrow \cos \alpha = \frac{1}{\sqrt{3}}$$

